**Informal lab write up. One per student. Name: \_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_ /20**

Partner's name(s) \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

***Due: Thursday, November 2nd or earlier***

Show all calculations. State answers in full sentences for full marks.

1. Collect all your data in a table. (Include mass of copper sulphate pentahydrate, filter paper & filter paper + precipitate. For full marks a table must have a title, border lines, units and be organized.) (3)
2. Write out the balanced reaction for this experiment. Include phase. (2)
3. How many moles of each reactant was used. (1)
4. Expected results.
	1. How many moles of each product was expected? Why? (2)
	2. Convert each product into grams. (2)
5. Calculate the % error using the expected mass and the mass found from the experiment. Show all calculations made. (2)

$$\% error= \frac{m\_{experimental}- m\_{calculated}}{m\_{calculated}}×100$$

1. Discussion:
	1. Did you end up with too much or too little precipitate? (1)
	2. Why do you think this happened? Explain. (Is this what you expected? Discuss reasons for this outcome.) (3)

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1. Suggest a way we could have found the mass of the Na2SO4 that was actually produced. (2)

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1. If you used 1.54 g of copper sulfate pentahydrate instead, how many grams of precipitate would be expected? (2)