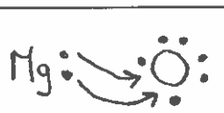
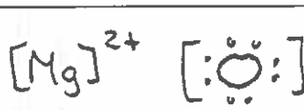
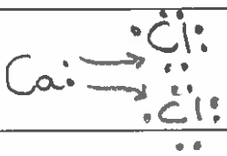
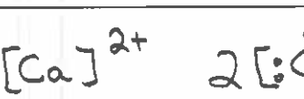
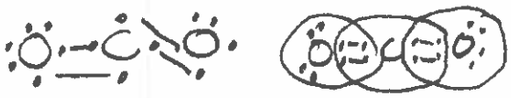
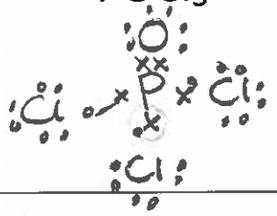
IN-CLASS SHEET finish for homework if necessary. **Name:** _____

The following compounds contain ionic bonds, they dissociate when dissolved in water (meaning they break apart into their ions). Draw the Lewis dot diagram for each atom before and after the dissociation as shown in your notes. (Electrons are transferred)

NaCl		
KI		
MgO		
$CaCl_2$		

Draw the Lewis dot diagrams for the following compounds. The following represent covalent bonds where the atoms share electrons! Please circle the octets.

H_2 	O_2 
CO_2 	$POCl_3$ 

Balance the following equations.

$$2\,H_2 + O_2 \rightarrow 2\,H_2O$$

$$2\,N_2 + 5\,O_2 \rightarrow 2\,N_2O_5$$

$$6\,K + B_2O_3 \rightarrow 3\,K_2O + 2\,B$$

$$2\,H_3PO_4 + 3\,Ca(OH)_2 \rightarrow Ca_3(PO_4)_2 + 6\,H_2O$$

Draw the Lewis Dot structure for the following polyatomic ions.

Ammonium NH_4^+ **Hydroxide OH^-** **Carbonate CO_3^{2-}**