Stoichiometry review sheet.

1. Name:
	1. KI \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	2. SnBr4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	3. CCl4  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. Cr(PO4)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

2. Determine the formula for:

* 1. lead (II) sulfate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	2. ammonium chromate \_\_\_\_\_\_\_\_\_\_\_\_\_
	3. potassium sulfide \_\_\_\_\_\_\_\_\_\_\_\_\_
	4. sodium hydroxide \_\_\_\_\_\_\_\_\_\_\_\_\_
1. Balance:
	1. C9H10O + O2 🡪 CO2 + H2O b. C4H10 + O2 🡪 CO2 + H2O
2. How many moles of sodium atoms correspond to 1.56x1021 atoms of sodium?
3. How many moles of H2 and N2 can be formed by the decomposition of 0.145 mol of ammonia, NH3?
4. The incandescent white of a fireworks display is caused by the reaction of phosphorous with O2 to give P4O10.
	1. Write the balanced chemical equation for the reaction: & write out the balance gram ratio below it:
	2. How many grams of O2 are needed to combine with 6.85g of P?

* 1. How many grams of P4O10 can be made from 8.00g of O2?
	2. How many grams of P are needed to make 7.46g P4O10?
1. Consider this neutralization reaction: **2 HCl + Mg(OH)2 → MgCl2 + 2 H2O**

**What mass of Mg(OH)2 is required to neutralize 4 moles of HCl?**